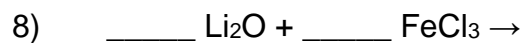
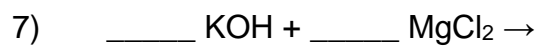
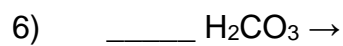
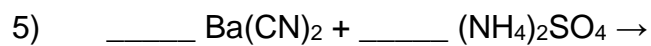
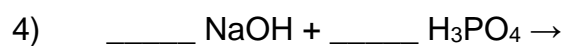
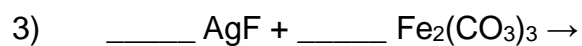
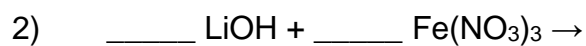
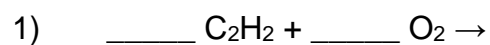


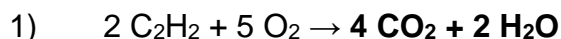
Predicting Reaction Products I

Predict the products of the following reactions and balance the resulting equations:

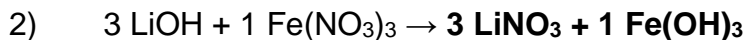


Predicting Reaction Products I – Answers

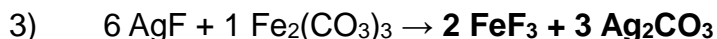
Predict the products of the following reactions and balance the resulting equations:



Like all combustion reactions, the products of this reaction are water and carbon dioxide.



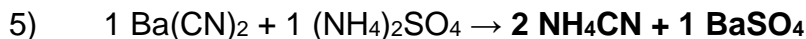
$\text{Fe}(\text{OH})_3$ is unstable in water, so the reaction will not occur.



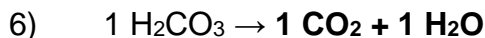
$\text{Fe}_2(\text{CO}_3)_3$ is unstable so the reaction won't occur.



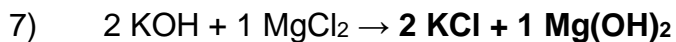
Acid-base reactions result in the formation of water and an ionic compound.



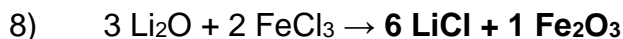
Because this is a double displacement reaction, the anions of both chemical compounds switch place.



There is only one compound in the reagent side of this equation, implying a decomposition reaction. When we perform decomposition reactions we try to find small compounds or elements that might result from this process, such as water and carbon dioxide.



Because this is a double displacement reaction, the anions of both chemical compounds switch place.



Because this is a double displacement reaction, the anions of both chemical compounds switch place.