

Chemistry  
Nomenclature & Oxidation Number Practice #1

Name: \_\_\_\_\_

For the following compounds, write the name and identify the oxidation states of each element:

1. CuF \_\_\_\_\_
2. CdI<sub>2</sub> \_\_\_\_\_
3. HI \_\_\_\_\_
4. NO \_\_\_\_\_
5. NF<sub>3</sub> \_\_\_\_\_
6. N<sub>2</sub>Cl<sub>2</sub> \_\_\_\_\_
7. CF<sub>4</sub> \_\_\_\_\_
8. KCl \_\_\_\_\_
9. NaNO<sub>3</sub> \_\_\_\_\_
10. Ca(NO<sub>2</sub>)<sub>2</sub> \_\_\_\_\_

For the following compounds, write the formulas and identify the oxidation states of each element:

1. Sulfur difluoride \_\_\_\_\_
2. Sulfur hexafluoride \_\_\_\_\_
3. Sodium phosphate \_\_\_\_\_
4. Lithium nitride \_\_\_\_\_
5. Chromium (III) carbonate \_\_\_\_\_
6. Tin (II) fluoride \_\_\_\_\_
7. Ammonium acetate \_\_\_\_\_
8. Ammonium hydrogen carbonate \_\_\_\_\_
9. Cobalt (III) nitrate \_\_\_\_\_
10. Copper (II) chloride \_\_\_\_\_