

## Naming Chemical Compounds Worksheet

Name the following ionic compounds:

1) NaBr \_\_\_\_\_

2) CaO \_\_\_\_\_

3) Li<sub>2</sub>S \_\_\_\_\_

4) MgBr<sub>2</sub> \_\_\_\_\_

5) Pu(OH)<sub>3</sub> \_\_\_\_\_

6) Hg(CN)<sub>2</sub> \_\_\_\_\_

7) Mo(C<sub>2</sub>H<sub>3</sub>O<sub>2</sub>) \_\_\_\_\_

8) Cr(CrO<sub>4</sub>)<sub>3</sub> \_\_\_\_\_

9) WO<sub>2</sub> \_\_\_\_\_

10) Mg(ClO)<sub>2</sub> \_\_\_\_\_

Write the formulas for the following ionic compounds:

11. potassium dichromate \_\_\_\_\_

12. gold(III) oxide \_\_\_\_\_

13. aluminum periodate \_\_\_\_\_

14. barium nitrite \_\_\_\_\_

15. silver carbonate \_\_\_\_\_

16. lithium sulfite \_\_\_\_\_

17. zinc hydrogen carbonate \_\_\_\_\_

18. iron(III) hydroxide \_\_\_\_\_

19. ammonium phosphate \_\_\_\_\_

20. copper(II) bromite \_\_\_\_\_

Name the following molecular compounds:

21) SO<sub>3</sub> \_\_\_\_\_

22) N<sub>2</sub>S \_\_\_\_\_

23) BF<sub>3</sub> \_\_\_\_\_

24) P<sub>2</sub>Br<sub>4</sub> \_\_\_\_\_

25) SiO<sub>2</sub> \_\_\_\_\_

26) SF<sub>6</sub> \_\_\_\_\_

27) NO<sub>2</sub> \_\_\_\_\_

Write the formulas for the following molecular compounds:

28) nitrogen trichloride \_\_\_\_\_

29) dinitrogen trioxide \_\_\_\_\_

30) phosphorus pentafluoride \_\_\_\_\_

31) diboron tetrahydride \_\_\_\_\_

Name the following acids and bases:

90) NaOH \_\_\_\_\_

91) H<sub>2</sub>SO<sub>3</sub> \_\_\_\_\_

92) H<sub>2</sub>S \_\_\_\_\_

95) NH<sub>3</sub> \_\_\_\_\_

96) HCN \_\_\_\_\_

97) Ca(OH)<sub>2</sub> \_\_\_\_\_

98) Fe(OH)<sub>3</sub> \_\_\_\_\_

Write the formulas of the following acids and bases:

99) hydrobromic acid \_\_\_\_\_

102) lithium hydroxide \_\_\_\_\_

103) nitrous acid \_\_\_\_\_

104) cobalt (II) hydroxide \_\_\_\_\_

105) sulfuric acid \_\_\_\_\_

106) beryllium hydroxide \_\_\_\_\_

Name the following chemical compounds

61) NaBr \_\_\_\_\_

62) Ca(C<sub>2</sub>H<sub>3</sub>O<sub>2</sub>)<sub>2</sub> \_\_\_\_\_

63) P<sub>2</sub>O<sub>5</sub> \_\_\_\_\_

64) Ti(SO<sub>4</sub>)<sub>2</sub> \_\_\_\_\_

65) FePO<sub>4</sub> \_\_\_\_\_

94) H<sub>3</sub>PO<sub>4</sub> \_\_\_\_\_

66) K<sub>3</sub>N \_\_\_\_\_

67) SO<sub>2</sub> \_\_\_\_\_

68) CuOH \_\_\_\_\_

93) H<sub>3</sub>P \_\_\_\_\_

69) Zn(NO<sub>2</sub>)<sub>2</sub> \_\_\_\_\_

70) V<sub>2</sub>S<sub>3</sub> \_\_\_\_\_

Write the formulas for the following chemical compounds:

71) silicon dioxide \_\_\_\_\_

72) nickel (III) sulfide \_\_\_\_\_

73) manganese (II) phosphate \_\_\_\_\_

101) carbonic acid \_\_\_\_\_

74) silver acetate \_\_\_\_\_

75) diboron tetrabromide \_\_\_\_\_

76) magnesium sulfate \_\_\_\_\_

77) potassium carbonate \_\_\_\_\_

100) hydrofluoric acid \_\_\_\_\_

78) ammonium oxide \_\_\_\_\_

79) tin (IV) selenide \_\_\_\_\_

80) carbon tetrachloride \_\_\_\_\_